Cambridge IGCSE™

CO-ORDINATED SCIENCES **0654/62**

Paper 6 Alternative to Practical **May/June 2022**

MARK SCHEME

Maximum Mark: 60

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2022 series for most Cambridge IGCSE, Cambridge International A and AS Level and Cambridge Pre-U components, and some Cambridge O Level components.

This document consists of **10** printed pages.

Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:   Marks must be awarded in line with: • the specific content of the mark scheme or the generic level descriptors for the question   • the specific skills defined in the mark scheme or in the generic level descriptors for the question   • the standard of response required by a candidate as exemplified by the standardisation scripts.
GENERIC MARKING PRINCIPLE 2:   Marks awarded are always **whole marks** (not half marks, or other fractions).
GENERIC MARKING PRINCIPLE 3:   Marks must be awarded **positively**:   • marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate   • marks are awarded when candidates clearly demonstrate what they know and can do   • marks are not deducted for errors   • marks are not deducted for omissions   • answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.
GENERIC MARKING PRINCIPLE 4:   Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

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GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1	Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
2	The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
3	Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
4	The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
5	'List rule' guidance For questions that require ***n*** responses (e.g. State **two** reasons …): • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked *ignore* in the mark scheme should not count towards ***n***. • Incorrect responses should not be awarded credit but will still count towards ***n***. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6   Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7   Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Question	Answer	Marks
1(a)	minutes ;	1
1(b)	3 of: iodine turns blue-black in presence of starch ; starch molecules are large; iodine molecules are small ; bag doesn't contain starch / starch cannot diffuse into the bag / starch cannot get though the membrane (because iodine brown / not blue-black) OR iodine (molecules move into) in beaker / starch / iodine (diffuses) out of the bag / iodine goes through the membrane (because iodine is blue-black) ;	3
1(c)	more precise ;	1
1(d)	avoid iodine on (outside) of the bag;	1
1(e)	longer time / slower to turn blue-black / change colour / more negative / brown results ;	1
1(f)(i)	Benedict's; heat; red / orange / yellow / green;	3
1(f)(ii)	biuret; lilac / purple / mauve / violet ;	2

Question	Answer	Marks
2(a)	clear and continuous outline; size $\geqslant$ ½ box ; 1 cell surrounded by 5 cells in correct shape ;	3
2(b)(i)	76 **and** line drawn;	1
2(b)(ii)	101 / 101.3 ;	1
2(c)	bulge on photo / thickness of the leaf varies AW; measure at several / many places; calculate average / mean;	3

Question	Answer	Marks
3(a)(i)	(gas) syringe ;	1
3(a)(ii)	burette / volumetric pipette ;	1
3(a)(iii)	stops fizzing / no (more) gas collected / formed ;	1
3(b)	solid / calcium carbonate **and** left (in the conical flask / at the end) ;	1
3(c)	34 ; 47 ;	2
3(d)(i)	axes correct way round and labelled with quantity and unit ; axes linear and plotted points cover $\geqslant$ ½ the grid ; points plotted correctly ;	3

Question	Answer	Marks
3(d)(ii)	point at 180 s ;	**1**
3(d)(iii)	line from origin showing curve and then level ;	**1**
3(e)(i)	any 2 from: reaction is fast at the start as the graph is steep ; rate of reaction decreases / reaction slowing down as the curve gets less steep ; reaction stops and line is horizontal / reaction stops and volume stays the same ;	**2**
3(e)(ii)	line drawn to the left of the original curve and doesn't cross it at any point before it goes horizontal ; to same height ;	**2**

Question	Answer	Marks
4	<table><tr><td rowspan="2">Test</td><td colspan="2">observations</td></tr><tr><td>calcium chloride</td><td>copper sulfate</td></tr><tr><td>add a few drops of aqueous ammonia add excess aqueous ammonia</td><td>white ppt ; (white ppt)</td><td>blue ppt **and** dark blue solution / dissolves to give dark blue colour ;</td></tr><tr><td>add a few drops of aqueous sodium hydroxide</td><td>white ppt **and**</td><td>blue ppt ;</td></tr><tr><td>add dilute nitric acid and aqueous barium nitrate</td><td>no ppt / colourless (solution) / no change / no reaction/ nothing happens **and**</td><td>white ppt ;</td></tr><tr><td>add dilute nitric acid and aqueous silver nitrate</td><td>white ppt **and**</td><td>no ppt / blue solution / no change / no reaction / nothing happens ;</td></tr></table>	5

Question	Answer	Marks
5(a)	correct symbol ; —(V)— correct parallel connection ;	2
5(b)(i)	V = 2.9 ; I = 0.61 ;	2
5(b)(ii)	cell/power supply may run down / wires may heat up ;	1
5(c)(i)	4.8 / 5 ;	1

Question	Answer	Marks
5(c)(ii)	Ω / ohm ;	1
5(d)	12×1.5 times = 18 10% 18 is 1.8 / 10% gives 16.2 AND not within 10% 10% 14 is 1.4 / 10% gives 15.4 AND not within 10%	2
5(e)	check if the ammeter gives a reading ;	1
5(f)(i)	**3** lamps connected in parallel ; 3 lamps drawn and rest of circuit correct ;	2
5(f)(ii)	variable resistor ;	1

Question	Answer	Marks
6	**One mark from each section and any two others** **additional apparatus:** stop-watch / timer ; thermometer ; measuring cylinder ; **method:** pour **hot** water into beaker and allow to cool/leave **and** more than one size / surface area ; measure temperature at start and end / time for a specific T drop / measure temperature every minute / time period ; 5 different sizes / surface areas ; repeat for each container ; **table:** columns for temperature and time / surface area against time ; with units ; **control variables:** volume of hot water ; initial hot water temperature ; room temperature ; time (of temperature loss) / temperature decrease ; **processing and conclusion:** calculate average to recognise / exclude anomalies ; calculate rate by loss of temperature ÷ time ; calculate surface area by πr^2 ; plot a graph of temperature against time / rate for each beaker and compare / surface area against time / rate ; use results to see whether increasing surface area increases or decreases the (rate of) temperature loss ;	**7**